

## ALUMINIUM AND COBALT SUBSTITUTED LITHIUM FERRITE PREPARED BY STANDERD CERAMIC METHOD

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### ABSTRACT:

An Aluminium and Cobalt substituted Lithium hardferrite ( M- type ) samples with the general chemical formula  $\text{Li}_{0.5}\text{Fe}_{0.5+x}\text{Co}_x\text{Al}_{12-2x}\text{O}_{19}$  were synthesized using reacting oxide by high temperature solid state reaction technique. The structural characterization of compound has been carried out from X-Ray diffraction powder pattern. The compounds are in single hexagonal phase without traces of uncertainly ambiguous reflection. From XRD pattern lattice parameters has recorded with increasing doped aluminum element in the range from  $a = 5.807 \text{ \AA}$  to  $5.906 \text{ \AA}$  and  $c = 22.507 \text{ \AA}$  to  $22.585 \text{ \AA}$  pertaining the space group  $P6_3/mmc$  (No.194) conformed from JCPDS data. The mass density of the ferrites were found linearly varies and depends upon the mass and volume of sample.

The X-Ray density has depends upon the lattice constant and molecular weight of the compounds. The average particle size was also estimated. The compounds were studies magnetically by using Guoy's method in the temperature range 300 to 800 K, the result shows that the compounds are paramagnetic in nature. The Curie molar constant was work out.

**Key words:** Magnetoplumbite hexaferrites, Molar constant, Seebeck coefficient.etc.

### 1. INTRODUCTION:

Lithium hardferrites with general chemical formula  $\text{Li}_{0.5}\text{Fe}_{12.5}\text{O}_{19}$  has been of great technological interest in many electromagnetic devices for a long time. High electrical resistivity, low eddy current losses, low magnetic losses, and very good thermal and chemical stability make lithium ferrite of a material of great importance for microwave applications. In the family of hexagonal ferrites, the mangnetoplumbite hexaferrites (M-Type) created much attention due to wide range of application in industries and created potential to interest in technological and scientific research due to their application importance such as high density magnetic recording, microwave device materials, hard disc in computer system. The application need particularly magnetic and electrical specification with the view, many attempt have been improved the properties of hexagonal ferrites using different tract of additives. The calcium hexferrites [1-6] have magnetic properties comparable to BaM and Sr M. In Calcium ferrites many attempt has been made to replace  $\text{Fe}^{+3}$  ions with  $\text{Al}^{+3}$ ,  $\text{Cr}^{+3}$  and  $\text{Co}^{+3}$  etc. A compound with the combination of bivalent-tetravalent cation was also used to replace  $\text{Fe}^{+3}$  ion such

as Cu-Ti, Co-Ti, Co-Sn, Zn-Sn etc[7, 8] without any appreciable change in BaM structure. When  $\text{Fe}^{+3}$  ions are replaced by non magnetic ions like  $\text{Ti}^{+4}$  and  $\text{Sn}^{+4}$  etc. In the same way Lithium ferrites substituted with Al and Co has been studied structurally, electrically and magnetically [9]. In the present study a series of five sample with chemical formula  $\text{Li}_{0.5}\text{Fe}_{0.5+x}\text{Al}_{12-2x}\text{Co}_x\text{O}_{19}$  ( $x = 2, 3, 4, 5$  and  $6$ ) were prepared. Lithium ferrites have attracted considerable attention because of the squareness of hysteresis loop coupled with superior temperature performance the crystal structure of M-type like compound with a space group  $\text{P6}_3/\text{mmc}$  (194) can be described as superposition of two structure block namely R-block with composition  $\text{BaFe}_6\text{O}_{11}$  and S-block with composition  $\text{Fe}_6\text{O}_8$  [10] in the stoichiometric ratio.

## 2. EXPERIMENTAL:

Multicrystalline sample powder were synthesized by high temperature solid state reaction using A.R grade oxides with stoichiometric ratio  $\text{Li}_2\text{O}$ ,  $\text{Fe}_2\text{O}_3$ ,  $\text{Al}_2\text{O}_3$  and  $\text{Co}_2\text{O}_3$  mixture.  $\text{Li}_2\text{O}$  oxide was carefully dehydrated before the mixing procedure. After grinding the mixture under acetone for six hours. During the preparation of pellets of thoroughly grounded mixture in the proper molar ratio with 5% of PVA as a binder are prepared by applying 10 tone pressure per square inch. These pellets slowly heated in the furnace at  $600^\circ\text{C}$  for 5 - 6 hours to remove binder. Then it was fired at  $1200^\circ\text{C}$  for 120 hours continuously, after that the furnace were cooled at the rate of  $20^\circ\text{C}$  per hours up to  $1000^\circ\text{C}$  and then  $50^\circ\text{C}$  per hours and then cooled in natural way the phase of final sample were verified

by Philips X-ray diffractometry using Ni-filter copper radiation. The mean grain size of multicrystalline samples was in the range  $150 - 200 \text{ \AA}$ . All the samples show single phase formation with a space group  $\text{P6}_3/\text{mmc}$  (194) confirmed from JCPDS data. X-ray pattern of the samples as shown in fig.1.1

The D.C resistivity or electrical conductivity was measured by the methods using LCR meter [11]. The end faces of pellets were coated with thin layer of conducting silver paste and measurement were made from room temperature to  $800 \text{ K}$ . Thermoelectric power measurement were carried out after sandwiching the thick pellet between two copper rods from room temperature to  $550^\circ\text{C}$

## 3. RESULT AND DISCUSSION:

In the present work, the Aluminium and Cobalt substituted Lithium hardferrites were introduced with general formula  $\text{Li}_{0.5}\text{Fe}_{0.5+x}\text{Al}_{12-2x}\text{Co}_x\text{O}_{19}$  ( $x = 2, 3, 4, 5$  and  $6$ ). The ions in Ba-M compounds can be replaced partly by  $\text{Co}^{+3}$  or completely  $\text{Li}^{+1}$  and combination of  $\text{Fe}^{+3}$  and  $\text{Al}^{+3}$  ions without changing the crystal lattice symmetry [12]. In all the specimens substituted ions would be chosen to keep electrical neutrality and to have a similar ionic radii in these ferrites. Due to the resemblance of ionic radii of  $\text{Fe}^{+3}$  with  $\text{Co}^{+3}$  and  $\text{Al}^{+3}$  ions, the ferrite ions will be replaced by cobalt and aluminium. It is seen that former ions are very easily replaced at any substituted variation in all specimens. The hexagonal lattice parameters 'a' and 'c' decreases linearly with the substitution variation  $\text{Co}^{+3}$  and  $\text{Fe}^{+3}$  concentration in all specimens. Thus the doping ions Cobalt

and Aluminium play an important role in the property variation of the compounds. XRD technique is used to confirm the formation of hexagonal M structure of compounds belonging to a space group  $P6_3/mmc$  (194) confirm from JCPDS data in

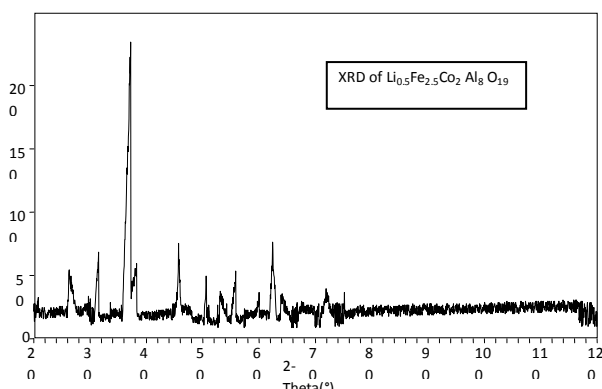


Fig 1.1

International Table for X-ray diffraction photograph Homawalt 1956 [13-16]. The decrease in lattice parameter and cell volume agree with result for Ba /Sr ferrite [17,18]. The numerical values of compositional data such as lattice constant, cell volume and X- ray density are tabulated in table -1. The observed value of electrical conductivity, activation energy and curie molar constant for specimens are also tabulated in table.2.

The plot of  $\ln \sigma$  vs  $(1/T) \times 10^{-3} K$  for the entire sample was almost linear. The electrical conductivity of these ferrites increases with increasing ferrite ion concentration. The electrical conductivity of sintered specimens varies from  $2.193 \times 10^{12} \Omega^{-1} cm$  to  $5.78 \times 10^{-6} \Omega^{-1} cm$  of these ferrites. The other workers have obtained a conductivity value of  $2 \times 10^{-2} \Omega^{-1} cm$  for Li-

ferrite that obtained is  $2.3 \times 10^{-6} \Omega^{-1} cm$  [19]

In the present work the electrical conductivity value obtained for the compounds is  $2.193 \times 10^{-12}$  to  $5.78 \times 10^{-6} \Omega^{-1} cm$ . The value of the electrical conductivity may be partly attributed to the low evaporation of lithium from the sample ,while the electrical conductivity values obtained by Venugopal Reddy et al 1984 is  $2.3 \times 10^{-6} (\Omega^{-1} cm)$  and  $2 \times 10^{-2} (\Omega^{-1} cm)$  Rezlescu N. et.al 1974 for lithium ferrite[20-22].The variation of activation energy with the substitutional parameters  $x$ , it is explain on the basis of Vewrway model [23], A small number of ferrous ions ( $Fe^{+2}$ ) are generally developed during sintering process which lead the conductivity in ferrites suggesting the hopping mechanism according ( $Fe^{+2} = Fe^{+3} + e^{-1}$ ) [24]. However these transition take place for a very small interval of time and are not detectable by the ordinary method.This valence exchange mechanism of Verwey may be considered for these ferrites as general and applicable to M-type ferrite.

Table 1:

Structural data of lithium hexaferrites

Compounds	Lattice parameters		Cell Volume ( $\text{\AA}^3$ )	Mol. Wt gm	X-Ray Density gm/cc
	a ( $\text{\AA}$ )	c ( $\text{\AA}$ )			
$Li_{0.5}Fe_{2.5}Co_2Al_8O_{19}$	5.859	22.323	663.81	780.89	3.9065
$Li_{0.5}Fe_{6.5}Co_6O_{19}$	5.695	21.262	597.22	1024.4	5.6963

Table 2 :  
Electrical conductivity of compounds

Compounds	Electrical Resistivity at room temp. $\Omega \text{ cm}^{-1}$	Activation energy E in (eV)	Electrical Conductivity at roomtemp.. $\Omega^{-1} \text{ cm}$
$\text{Li}_{0.5}\text{Fe}_{2.5}\text{Al}_8\text{Co}_2\text{O}_{19}$	$4.559 \times 10^{11}$	0.77	$2.193 \times 10^{-12}$
$\text{Li}_{0.5}\text{Fe}_{6.5}\text{Co}_6\text{O}_{19}$	$1.72 \times 10^5$	0.39	$5.78 \times 10^{-6}$

#### 4. CONCLUSION:

In this present work of lithium hard ferrites is to check the formation of ferrite containing  $\text{Al}^{+3}$  and  $\text{Co}^{+3}$  ions along with  $\text{Fe}^{+3}$  ions. The values of lattice parameters  $a$  and  $c$  confirms the formation of hexagonal unit cell. All these compounds have M-structure through the site distribution changes. No changes occur in the charge distribution, but the site distribution is change due to strichiometric changes in the compound. In this series of  $\text{Li}_{0.5}\text{Fe}_{6.5}\text{Co}_6\text{O}_{19}$  to  $\text{Li}_{0.5}\text{Fe}_{2.5}\text{Co}_2\text{Al}_8\text{O}_{19}$  the electrical conductivity of these ferrites increases with increasing ferrite ion concentration. Also, by carried out the magnetic study, the Coerecivity, Retentivity, Saturation magnetization, magnetic moment decreased with Al ions substitution. It is evident from the fact that Al ions is weak magnetic in nature. Lithium hard ferrites shows magnetic properties which are more useful in the field of information storage recording media.

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#### REFERENCES:

1. J.Beretea and T.Brown. Austral. J.Chem.24(1971) 273
2. J.Lipka,A.Gruskova,O.Orlicky, J.Siteck, M.Miglierini,R.Grone, M.Hud and Toth Hyperfine Interaction 59(1990) 381.
3. G.Albanese, A Deriu, E.Luchini and G.Slokar.Appli Phys A 2(1981) 45
4. R.Muller, H.Pfeiffer and W.Schuppert, J.Mogn.Mogn Mater 18 (1991)101
5. D.K.Kulkarni and C.S.Prakash Bull Mater Sci.17(1994)
6. 35.G.Asti. M.Carbuccocchio. A Deriu, E.Lucchini and G.Slokar, J.Mogn mogn Mater 20(1980)44
7. J.G.Renson,J.A.Schullces and J.S.Van. Wlering,J.Phys Collog 32(1971)C1-924.
8. B.X.Gu, H.Y.Zang. H.R.Zhai, B.G.Slen, M.Lu, S.Y.Zhang and Y.Z.Maoi, Phys State Sol 133 (1992) K83
9. X.Obrador, A Isalgue,A Collomb, A Tejada, J.C.Joubert, J.,PhysC. 19(1986) 6605
10. D.B.Ghare, A.P.B.Sinha, J.Phys Chem Solid 29(1958) 885
11. K.Haneda, H.Kojima. Phys State Solid (A) 6(1971) 256
12. Kanke E, Takayama Muromachi Y.Uchida, Kato and S. Takikawa, J.Solid state Chem 95(1991) 43
13. Hanmawalt (1936) International Table for X-ray diffraction photograph.
14. L.G.Van Uitert, J. Appl. Phys 28.1 (1957)317

15. A. Isalgue, A Laberta, J.Tejada, X Obradir. Appl.Phys A39(1986) 221
16. S.S.Darokar, K.G.Rewatkar & D.K. Kulkarni; Mater Chem.Phys.56 (1998) 84-85
17. K. Haneda, M. Kojima, J. Appl. Phys 14.B(1973) 3760
18. V. Adelskod, Arkir Kemi Min Geo 12 A (1938)1
19. Albense G, Carbulichhio and Deril A J.Phys Solid State: A 23 (1974) 351
20. Venugopal Reddy et.al, J. Pure and Applied Physics 22,(1984)596,
21. N.Rezlescu, D.Condurachi, P.Petrarju and E.Lucca, J.A.M .Ceream Soc.57 (1974)40
22. S.S.Darokar, K.G.Rewatkar, M.S. Chowkase and D.K.Kulkarni, Indian J. Phys 74A(2) (2000) 155-157
23. E.J. Verway J.H, De-Bar.Pec.Trar.Chem Pay Bull 55 (1936)531
24. S.S. Darokar et.al, J. Adv. Appl. Sci. Research, 2013, 4(1): 173-177.